

100  $\mu$ l at 10ml level in 0.05 N HCl

$$19,731 \text{ counts} / 4 \text{ min} = 4932.8 \text{ cpm}$$

$$\Rightarrow 4.93 \times 10^5 \text{ counts in entire sample (10ml)}$$

Since this is JE-4,

$$\frac{4.93 \times 10^5 \text{ cpm}}{2858 \frac{\text{cpm}}{\mu\text{g}}} = 173 \mu\text{g adduct}$$

this is low  $\tilde{\times} 10$  - ? 90% quenching

It is possible that much of the  $^3\text{H}$  exchanged out of Adduct when Al Nadram exchanged protons with deuterium.

---

Subject

Instructor's Name

Preparative Run on Adduct used for NMR Work

12-14-76

On 12/13, Paul dissolved the NMR adduct sample in 10 ml. 0.05 N HCl - it dissolved well.

- inject 50  $\mu$ l 12-13-1 P - basically, the sample is pure, but there was a fair amount of diol (?) formed as a degradation product. It was fluorescent.
- 12/14 neutralize sample with 5 ml 0.1 N KOH
- Sample was left  $\sim$  pH 5.

12-14-76 Fluorescence Profile of 12-7-5 (Mod. DNA Hydrolysate)

- inj. 50  $\mu$ l 12-7-5
- sens 970 = 0.02

Results: 4 fluor. peaks, one eluting at  $t_R$  <sup>close to</sup>  $t_R$  of adduct.